SCIENTIFIC PROCESS



Water Molecules & Surface Tension

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Everything is made up of **MOLECULES**.



Water is made up of molecules.





The water's surface is CONVEX.

The molecules that make up water are constantly moving.

They are pushing and bouncing all around.



When matter is heated, the motion of its molecules increases.



The hotter the H₂O, the faster the water molecules move around.



H₂O molecules cling together to keep from separating.



This creates a bulge of H₂O on top.

The water's molecules are very attracted to each other. This causes SURFACE TENSION.





Surface temsion can be described as the water's SKIN. This is how certain bugs walk on water. That is why water forms as a drop.



When the surface tension has been stretched out as far as it can, the H2O breaks the skin and spills over.



Soap weakens water molecules.





When water molecules are attracted to each other, it is called COHESION.





When water molecules are attracted to the container they are in, it is called ADHESION.

Soapy water molecules are not very cohesive.

Rubbing alcohol molecules are not very cohesive.



Water molecules are **VERY** cohesive.



The surface tension, or cohesive force, is weakened by soap at that spot. The molecules that are further away from this spot will pull these weaker molecules towards them because they are still strong and the pepper & toothpicks will go with them.

THE SCIENTIFIC PROCESS



Step I: Scientific Process

Procedures

What **steps** are involved in doing this experiment? What will the **directions** be?



Step 2: Scientific Process

Materials

What **supplies** are needed to do this experiment?



Step 3: Scientific Process

Prediction

What do I **think** will happen in this experiment? What will be my hypothesis?



Step 4: Scientific Process



How did my experiment work?What happened & why?

